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Chapter 7 Chemistry Study Guide. STUDY. PLAY. valence electrons. electrons in the highest occupied energy level of an element's atoms. Group Number, 1A=1, 2A=2, 3A=3, 4A=4, 5A=5, 6A=6, 7A=7, 8A=8. How can you determine the number of valence electrons an atom has from its position on the periodic table? Name by that

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Chemistry Chapter 7 Study Guide. STUDY. PLAY. Anion. An ion with a negative charge. The charge for an Anion is written as a number followed by a minus sign. Cation. An ion with a positive charge. The charge for a cation is written as a number followed by a plus sign. Chemical formula.

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chemistry study guide chapter 7. what follows an element's symbol in a chemical formula to indicate the number of atoms of that element in one molecule of the compound? what does the chemical formula for an ionic compound represent?

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Chapter 7 Chemistry Study Guide. STUDY. PLAY. Explain why most carbon compounds are classified as organic compounds. Carbon compounds controls what happens in the cell of living organism and play key roles in the chemistry of living organisms. Why is lead no longer used in paints for plumbing pipes.

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Chemistry Chapter 7 Study Guide. Define chemical bond. Explain formation of cations and anions. Draw electron dot diagrams. Write electron configurations for ions. The force that holds two atoms together. Cations form when an atom loses electrons. An anion forms when.... Ex. K->1s^2 2s^2 2p^6 3s^2 3p^6 4s^1.

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Chemistry Chapter 7 Study Guide. the force that holds two atoms together; may form by the attraction of a positive ion for a negative ion or by sharing electrons. a three-dimensional geometric arrangement of particles in which each positive ion is surrounded by negative ions and each negative ion is surrounded by positive ions; vary in shape due to sizes and relative numbers of the ions bonded.

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