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## CHAPTER 18: Chemical Equilibrium I.

a. Write the chemical equilibrium expression for the following equations. Include the value of  $K$ . (a)  $(g) K=0.1$  (b)  $(g) + O_2(g) K = 1.8 \times 10^{-2}$  - Does the reaction favor products or reactants? (Will there be mostly products or reactants when it reaches

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2. a. Compare the rates of forward and reverse reactions when equilibrium has been reached. b.

North St. Paul-Maplewood Oakdale /  
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Equilibrium, equilibrium constant,  
chemical equilibrium expression, a  
chemical reaction in which the  
products can react to re-form....  
when the rate of its forward chemical  
reaction equals the rate....

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18-1 OBJECTIVES Define chemical  
equilibrium. Explain the nature of the  
equilibrium constant. Write chemical  
equilibrium expressions and carry out

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calculations involving them. FIGURE  
18-1 When heated, mercury(II) oxide  
decomposes into the elements from  
which it was

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558 Chapter 18 Chemical Equilibrium  
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will discover that many reactions and processes reach a state of equilibrium. You will use Le Châtelier ' s principle to explain how various factors affect chemical equilibria. You will calculate equilibrium concentrations of reactants and products using the equilibrium

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Equilibrium. Equilibrium. Value of  $K$ .  
 $K_{eq} < 1$ .  $K_{eq} > 1$ . when the rates of  
opposing actions are the same.  
represents the reaction progression...

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-all  $K_{eq}$ 's are dependent on...

reactants of the reaction are favored.

products of the reaction are favored.

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reaction in which the products can react to reform.... When the rate of its forward reaction equals the rate of its r.... The ratio of the mathematical product of the concentrations of.... A chemical reaction in which the products can react to reform....

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a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products

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occurs in the chemical system (18.2)  
equilibrium position the relative  
concentrations of reactants and  
products of a reaction that has  
reached equilibrium; indicates  
whether the reactants or products are  
favored in the reversible reaction  
(18.2)

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SHORT ANSWER Answer the following  
questions in the space provided. 1.

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Match the solution type on the right to the corresponding relationship between the ion product and the  $K_{sp}$  for that solution, listed on the left.

\_\_\_\_\_ The ion product exceeds the  $K_{sp}$ .

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### Equilibrium Answers

1. In a bottle of unopened cola, the  $\text{CO}_2$  gas dissolved in the liquid is in equilibrium with the  $\text{CO}_2$  gas above the liquid. The dissolved gas reacts with water molecules in the cola to form carbonic acid, which also dissociates into carbon dioxide and

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water. Which chemical equation(s)  
best describe this equilibrium  
system? a.  $\text{CO}_2(\text{g}) \rightleftharpoons \text{CO}_2(\text{l})$  b.

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between the ion product and the  $K_{sp}$

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for that solution, listed on the left.

The ion product exceeds the  $K_{sp}$ .  
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